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Exercise $\backslash(\backslash\text{PageIndex}\{2\}\backslash)$ We can determine the relative acid strengths of $\backslash(\backslash\text{ce}\{\text{NH}_4^+\}\backslash)$ and HCN by comparing their ionization constants. The ionization

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constant of HCN is given in Table E1 as 4.9×10^{-10} . The ionization constant of NH_4^+ is not listed, but the ionization constant of its conjugate base, NH_3 , is listed as 1.8×10^{-5}

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